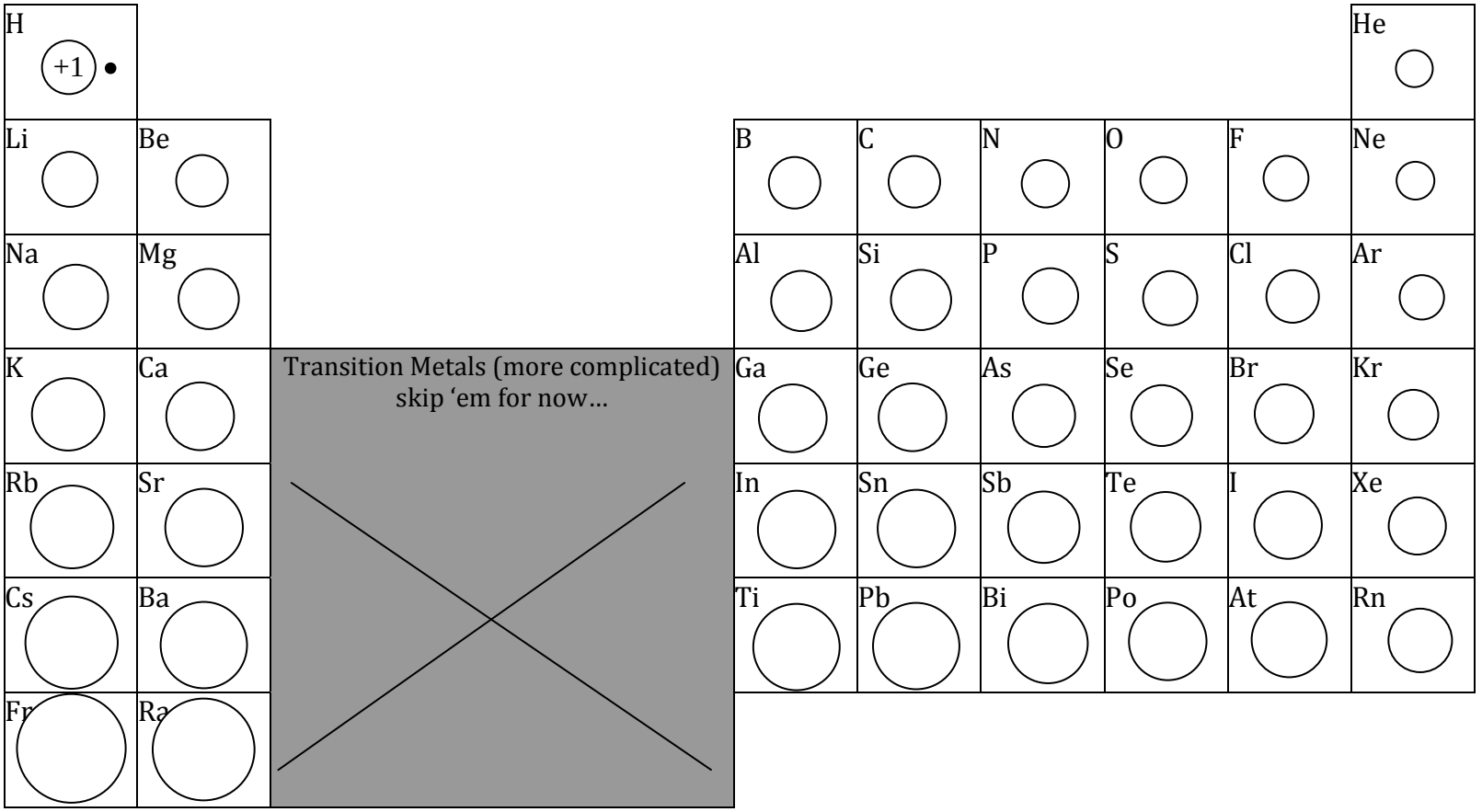
Periodic Trends

a) Fill in "core valence" charge for each element on the periodic table below. Show valence electrons as dots in the same configuration used for Lewis diagrams. *Hydrogen has been done for you as an example.*

b) Draw arrows on the top and to the left of the table to indicating the periodic trends for increasing atomic radius, electronegativity, electron affinity, and 1st ionization energy. Your teacher may instruct you to add more arrows for trends such as effective nuclear charge and reactivity.