

PART A: MULTIPLE CHOICE (10 MARKS)

1	2	3	4	5	6	7	8	9	10
c	a	b	d	d	a	b	b	c	a

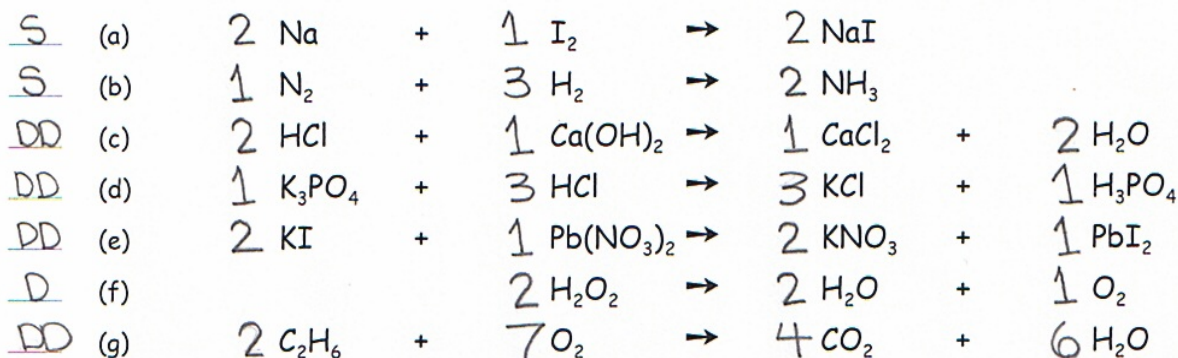
PART B: MATCH (5 MARKS)

1	2	3	4	5
H	E	J	C	F

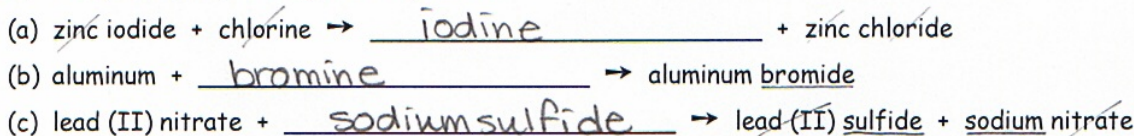
PART C: SHORT ANSWER (40 MARKS)

Answer the following questions in the space provided. If more space is needed use the back of this sheet.

- {2} 1. Examine the following word equation: $\text{propane} + \text{oxygen} \rightarrow \text{carbon dioxide} + \text{water}$
- (a) List one reactant: propane, oxygen (b) List one product: carbon dioxide, water
- {7} 2. (i) Classify each of the chemical reactions below as either (S) synthesis, (D) decomposition, (SD) single displacement, or (DD) double displacement. Place your answer in the space provided.
- {12} (ii) Balance each of the reactions below. (Be sure to use the number 1 when only 1 molecule is needed!)

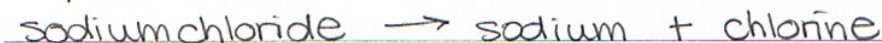


- {4} 3. Complete the following word equations.



4. Write word equations for each of the chemical reactions below.

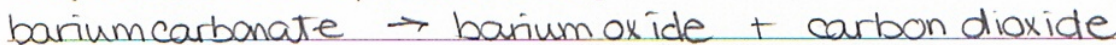
- {3} (a) NaCl produces Na and Cl



- {4} (b) Ca(OH)_2 and HBr react to form H_2O and CaBr_2



- {3} (c) BaCO_3 reacts when heated to produce BaO and CO_2



- {5} 5. Use the law of conservation of mass to answer the following:

(a) When 74.1 g of calcium is reacted with oxygen, 75.6 g of calcium oxide is produced. What mass of oxygen was used in the reaction?

(a) 1.5 g

(b) 36.1 g of sulfur reacts with 13.0 g of copper to produce copper(II) sulfide. What is the mass of copper(II) sulfide produced?

(b) 49.1 g

(c) What law did you use to answer these questions?

(c) conservation of mass